

Unit 8 Review

1. The balanced equation  $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$  tells us that 4.0 mol  $\text{H}_2$   
 A) reacts with 2.0 mol  $\text{P}_4$   
 B) produces 8.0 mol  $\text{PH}_3$   
 C) cannot react with phosphorus  
 D) produces 2.7 mol  $\text{PH}_3$   
 E) reacts with 4.0 mol  $\text{P}_4$

2. A 1.0-mol sample of  $\text{Fe}_2\text{O}_3$  was decomposed according to the equation:  
 $2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}_2\text{O}(\text{s}) + 3\text{O}_2(\text{g})$

How many moles of  $\text{O}_2$  are formed assuming 100% yield?

- A) 0.9 mol
- B) 1.5 mol
- C) 1.2 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ . If you have a 1.00-mol sample of iron, how many moles of  $\text{Fe}_2\text{O}_3$  will there be after the iron has rusted completely?

- A) 0.527 mol
- B) 0.750 mol
- C) 1.00 mol
- D) 1.50 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.  
 $8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_2\text{X}_4(\text{s}) + 3\text{O}_2(\text{g})$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of  $\text{B}_2\text{X}_4$ , what is the identity of X?

- A) Ca
- B) Mg
- C) Si
- D) Sn
- E) C

5. Refer to the following equation:  
 $4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$

How many molecules of  $\text{NO}$  are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B)  $3.84 \times 10^{23}$
- C)  $1.82 \times 10^{23}$
- D) 1.39
- E) none of these

6. In the reaction  $2\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ , how many moles of ammonia would be produced from 1.00 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.37 mol
- C) 0.727 mol
- D) 2.00 mol
- E) 0.500 mol

7. Refer to the following unbalanced equation:  
 $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

What mass of oxygen ( $\text{O}_2$ ) is required to react completely with 13.1 g of  $\text{C}_2\text{H}_6$ ?

- A)  $2.99 \times 10^2$  g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A)  $1.37 \times 10^{23}$  molecules
- B)  $1.00 \times 10^{23}$  molecules
- C)  $4.14 \times 10^{23}$  molecules
- D)  $3.10 \times 10^{23}$  molecules
- E)  $5.17 \times 10^{23}$  molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction.



- A) 79.2 g  $\text{CO}_2$
- B) 158 g  $\text{CO}_2$
- C) 89.1 g  $\text{CO}_2$
- D) 1.258 g  $\text{CO}_2$
- E) 1.400 g  $\text{CO}_2$

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A)  $1.03 \times 10^{23}$  molecules  $\text{O}_2$
- B)  $2.28 \times 10^{23}$  molecules  $\text{O}_2$
- C)  $1.71 \times 10^{23}$  molecules  $\text{O}_2$
- D)  $3.42 \times 10^{23}$  molecules  $\text{O}_2$
- E)  $5.13 \times 10^{23}$  molecules  $\text{O}_2$

# Multiple Choice Stoichiometry Test

**RJ Shavelson**



**Multiple Choice Stoichiometry Test:**

The Top Books of the Year Multiple Choice Stoichiometry Test The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Lets delve into the realm of top-selling books, exploring the captivating narratives that have charmed audiences this year. The Must-Read : Colleen Hoover's "It Ends with Us" This touching tale of love, loss, and resilience has gripped readers with its raw and emotional exploration of domestic abuse. Hoover masterfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can succeed. Uncover the Best : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This intriguing historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids absorbing storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Discover the Magic : Delia Owens "Where the Crawdads Sing" This captivating coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens crafts a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a exceptional and gripping novel that will keep you wondering until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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