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web oxidation reduction redox reactions oxidation $\text{Cu} + \text{Cu}^{2+} + 2\text{e}^-$ reduction $\text{MnO}_4^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$
l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} + \text{MnO}_4^- + 4\text{H}^+$ is

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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transformer at the sub station and transmit it with the 11 kv feeder to the village or

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