

## Chapter 9: Chemical Quantities

1. Solve stoichiometry problems using grams  $\rightarrow$  moles  $\rightarrow$  moles  $\rightarrow$  grams calculations. Examples below:

a. Calculate moles/grams of  $(\text{NH}_4)_2\text{SO}_4$  produced when 30.0 moles of  $\text{NH}_3$  react with  $\text{H}_2\text{SO}_4$ .



30.0 mol x moles

$$30.0 \text{ moles} \times 1 \text{ mol}/2 \text{ mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ grams/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

b.  $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$ . How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams  $\rightarrow$  Moles  $\rightarrow$  Moles  $\rightarrow$  Grams

$$100. \text{ g Na}_2\text{CO}_3 \times 1 \text{ mol}/106.0 \text{ g} \times 2 \text{ mol NaNO}_3 / 1 \text{ mol Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g NaNO}_3 \times 1 \text{ mol}/85.00 \text{ g} \times 2 \text{ mol} / 2 \text{ mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.  
What is the theoretical yield of  $\text{NaNO}_3$  if 100.0 grams of  $\text{Na}_2\text{CO}_3$  are used in the reaction? What is the percent yield of  $\text{NaNO}_3$  if 130.5 grams of  $\text{NaNO}_3$  are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield}/\text{Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? **The reactant that is used up first, thereby limiting the amount of product that can be formed.**

## Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

**Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.**

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? **Absorb a photon of energy.**

**How do they lose energy to return to their ground state? Release a photon of same energy as that absorbed.**

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? **Specific wavelengths of light seen in emission spectra for atoms (page 285-86)**

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

**Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.**

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

**See book, notes, probs 54, 96**

# 2nd Semester Chemistry Exam Review Fot Plato

**RJ Shavelson**



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