

Unit 8 Review

1. The balanced equation $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$ tells us that 4.0 mol H_2
 A) reacts with 2.0 mol P_4
 B) produces 8.0 mol PH_3
 C) cannot react with phosphorus
 D) produces 2.7 mol PH_3
 E) reacts with 4.0 mol P_4

2. A 1.5-mol sample of Fe_2O_3 was decomposed according to the equation:
 $2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}_2\text{O}(\text{s}) + 3\text{O}_2(\text{g})$

How many moles of O_2 are formed assuming 100% yield?
 A) 0.9 mol
 B) 1.5 mol
 C) 2.2 mol
 D) 0.7 mol
 E) 2.0 mol

3. The rusting of iron is represented by the equation $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. If you have a 1.04-mol sample of iron, how many moles of Fe_2O_3 will there be after the iron has rusted completely?

- A) 0.527 mol
- B) 0.790 mol
- C) 1.05 mol
- D) 1.58 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.
 $8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_2\text{X}_6(\text{s}) + 3\text{X}_2\text{O}_3(\text{g})$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of B_2X_6 , what is the identity of X?

- A) Ge
- B) Mg
- C) Si
- D) Sn
- E) C

5. Refer to the following equation:
 $4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$

How many molecules of NO are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B) 3.84×10^{23}
- C) 1.82×10^{23}
- D) 1.39
- E) none of these

6. In the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$, how many moles of ammonia would be produced from 1.09 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.27 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.363 mol

7. Refer to the following unbalanced equation:
 $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 What mass of oxygen (O_2) is required to react completely with 13.1 g of C_2H_6 ?

- A) 2.99×10^2 g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A) 1.37×10^{23} molecules
- B) 1.03×10^{23} molecules
- C) 4.14×10^{23} molecules
- D) 3.10×10^{23} molecules
- E) 5.17×10^{23} molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane, C_8H_{18} , in the following reaction.



- A) 79.2 g CO_2
- B) 158 g CO_2
- C) 89.1 g CO_2
- D) 1.258 g CO_2
- E) 1.400 g CO_2

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A) 1.03×10^{23} molecules O_2
- B) 2.28×10^{23} molecules O_2
- C) 1.71×10^{23} molecules O_2
- D) 3.42×10^{23} molecules O_2
- E) 5.13×10^{23} molecules O_2

Multiple Choice Stoichiometry Test

Christian Drosten



Multiple Choice Stoichiometry Test:

The book delves into Multiple Choice Stoichiometry Test. Multiple Choice Stoichiometry Test is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Multiple Choice Stoichiometry Test, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Multiple Choice Stoichiometry Test
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 - Chapter 3: Multiple Choice Stoichiometry Test in Everyday Life
 - Chapter 4: Multiple Choice Stoichiometry Test in Specific Contexts
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