

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH<sup>-</sup>] and [H<sup>+</sup>].
3. Calculate the pH of weak (or strong) acids.

## Vocabulary

- |                   |                                |        |
|-------------------|--------------------------------|--------|
| 1. autoionization | 2. conjugate acid-base pair    | 3. pH  |
| 4. hydronium ion  | 5. hydronium ion concentration | 6. pOH |
|                   | 7. hydroxide ion               |        |

## Key Equations

- |   |                         |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$                             | 4. $pOH = -\log[OH^-]$  |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Complete the reaction.

1.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)
2.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)
3.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)
4.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)
5.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)
6.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)
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8.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)
9.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)
10.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$  (autoionization of water)

11.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

12.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

13.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

14.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

15.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

16.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

17.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

18.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

19.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

20.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

21.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

22.  $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

# 19 2 Hydrogen Ions And Acidity Work Sheet Answers

**Rachel Sandford**



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